

**Chemistry**  
**UNIT: 4CH1**  
**PAPER: 2CR**

Total Marks
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**Tuesday 11 June 2024 – Morning**

**Time: 1 hour 15 minutes**

**In the boxes below, write your name, centre number and candidate number.**

<b>Surname</b>					
<b>Other names</b>					
<b>Centre Number</b>					
<b>Candidate Number</b>					

## **YOU MUST HAVE**

**Calculator, ruler**

## **YOU WILL BE GIVEN**

**Periodic Table**

**Diagram Booklet**

## **INSTRUCTIONS**

**Answer ALL questions.**

**Answer the questions in the spaces provided in this Question Paper or in the separate Diagram Booklet – there may be more space than you need.**

**Show all the steps in any calculations and state the units.**

## **INFORMATION**

**The total mark for this paper is 70**

**The marks for EACH question are shown in brackets – use this as a guide as to how much time to spend on each question.**

**There may be spare copies of some diagrams.**

**ADVICE**

**Read each question carefully before you start to answer it.**

**Try to answer every question.**

**Check your answers if you have time at the end.**

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**Answer ALL questions.**

**Some questions must be answered with a cross in a box ☐. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☐.**

**1. This question is about sulfur.**

**(a) Look at the diagram for Question 1 (a) in the Diagram Booklet. It represents the electronic configuration of an atom of sulfur.**

**(i) Which period of the Periodic Table contains sulfur?  
(1 mark)**

☐ **A 2**

☐ **B 3**

☐ **C 4**

☐ **D 6**

**(continued on the next page)**

1. (a) continued.

(ii) What is the charge on a sulfide ion?  
(1 mark)

☐ A 1+

☐ B 2+

☐ C 1–

☐ D 2–

(b) When a mixture of sulfur and zinc is heated to a high temperature a reaction occurs, forming the compound zinc sulfide, **ZnS**

(i) Give a reason why the mixture of sulfur and zinc needs heating before a reaction occurs.  
(1 mark)

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(continued on the next page)

1. (b) continued.

- (ii) Calculate the relative formula mass ( $M_r$ ) of zinc sulfide.  
(1 mark)

[for Zn,  $A_r = 65$       for S,  $A_r = 32$ ]

$M_r =$  \_\_\_\_\_

(Total for Question 1 = 4 marks)

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2 (a) Look at the diagram for Question 2 (a) in the Diagram Booklet. It represents an atom of an element.

(i) What is the name of the particle labelled **W**?  
(1 mark)

☐ A electron

☐ B ion

☐ C neutron

☐ D proton

(ii) What is the mass number of this atom?  
(1 mark)

☐ A 4

☐ B 5

☐ C 9

☐ D 13

(continued on the next page)

2. continued.

(b) These are the symbols for the two isotopes of lithium.



- (i) In terms of sub-atomic particles, give one similarity and one difference between the two isotopes.  
(2 marks)

similarity

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difference

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(continued on the next page)



2. (b) continued.

- (ii) A sample of lithium contains 7.5% of  ${}^6_3\text{Li}$   
and 92.5% of  ${}^7_3\text{Li}$

Calculate the relative atomic mass ( $A_r$ ) of  
this sample of lithium.  
(2 marks)

$A_r =$  \_\_\_\_\_

(Total for Question 2 = 6 marks)

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**3 This question is about magnesium metal and its compounds.**

**(a) (i) Look at the diagram for Question 3 (a) (i) in the Diagram Booklet. It represents the structure of magnesium.**

**Add the labels for A and B below.  
(2 marks)**

**A** \_\_\_\_\_

\_\_\_\_\_

**B** \_\_\_\_\_

\_\_\_\_\_

**(continued on the next page)**

**3. (a) continued.**

**(ii) Explain why magnesium is malleable.  
(2 marks)**

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**3. continued.**

**(b) Magnesium burns in oxygen gas to form solid magnesium oxide.**

**(i) Look at the diagram for Question 3 (b) (i) in the Diagram Booklet. Complete the chemical equation for the reaction by balancing the equation and adding the state symbols.  
(2 marks)**

**(ii) Give a reason why magnesium is oxidised in this reaction.  
(1 mark)**

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**(continued on the next page)**

3. continued.

(c) A student uses this method to make a solution of magnesium chloride.

Step 1 measure  $25\text{ cm}^3$  of dilute hydrochloric acid into a beaker

Step 2 add magnesium powder a little at a time

Step 3 keep adding magnesium powder until it is in excess

Step 4 remove the excess magnesium powder by filtration

This is the equation for the reaction.



(i) Give a reason why the student uses magnesium powder rather than magnesium ribbon.  
(1 mark)

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**3. (c) continued.**

- (ii) Give a reason why the student adds an excess of magnesium in step 3  
(1 mark)**

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**(continued on the next page)**

3. (c) continued.

- (iii) Calculate the minimum mass of magnesium needed to react with  $25.0 \text{ cm}^3$  of  $2.00 \text{ mol/dm}^3$  hydrochloric acid.  
(3 marks)

[for Mg,  $A_r = 24$ ]

mass = \_\_\_\_\_ g

(Total for Question 3 = 12 marks)

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**4. This question is about the manufacture of ethanol by the fermentation of glucose.**

**(a) Fermentation needs to be done in the absence of air.**

**Give TWO other conditions needed for fermentation.  
(2 marks)**

**1** \_\_\_\_\_

\_\_\_\_\_

**2** \_\_\_\_\_

\_\_\_\_\_

**(b) Explain why fermentation needs to be done in the absence of air.  
(2 marks)**

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_



4 . continued.

- (c) (i) Look at the diagram for Question 4 (c) (i) in the Diagram Booklet. Complete the equation for fermentation.

(1 mark)

- (ii) A mixture containing **500 mol** of glucose is fermented.

A mass of **5750 g** of ethanol is obtained.

Calculate the percentage yield.

(3 marks)

[for ethanol,  $M_r = 46$ ]

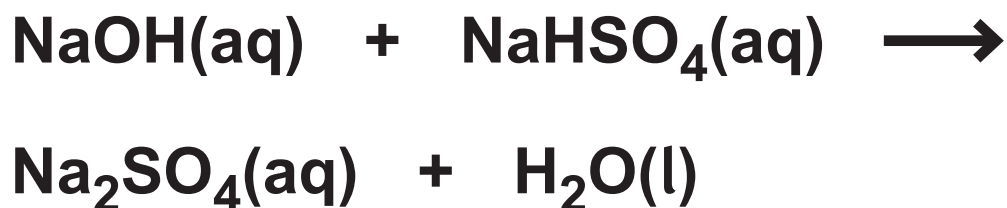
percentage yield = \_\_\_\_\_ %

(Total for Question 4 = 8 marks)

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5. Sodium sulfate can be prepared by the reaction between sodium hydroxide solution and sodium hydrogensulfate ( $\text{NaHSO}_4$ ) solution.

This is the equation for the reaction.



Sodium hydrogensulfate solution is acidic.

A student adds  $25.0 \text{ cm}^3$  of sodium hydroxide solution to a conical flask and adds two drops of indicator.

The student does a titration.

- (a) (i) Name a piece of apparatus that should be used to add  $25.0 \text{ cm}^3$  of sodium hydroxide solution to the conical flask.  
(1 mark)

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**5. (a) continued.**

- (ii) Describe the method the student should use to find the accurate volume of sodium hydrogensulfate solution needed to neutralise the  $25.0\text{ cm}^3$  of sodium hydroxide solution.**  
**(5 marks)**

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**5. (a) (ii) continued.**

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**5. continued.**

- (b) The student repeats the titration without the indicator and forms a solution of sodium sulfate.**

**Describe how the student can obtain pure, dry crystals of hydrated sodium sulfate from the solution.**

**(4 marks)**

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**Turn over**

**5. (b) continued.**

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5. continued.

- (c) Crystals of hydrated sodium sulfate decompose when heated.

This is the equation for the decomposition.



Look at the diagram for Question 5 (c) in the Diagram Booklet. A student uses the apparatus to find the value of **X**.

The student heats the crystals until the decomposition is complete.

Look at the table for Question 5 (c) in the Diagram Booklet. It shows the student's results.

Use the results to calculate the value of **X**.  
(4 marks)

[for  $\text{Na}_2\text{SO}_4$ ,  $M_r = 142$

for  $\text{H}_2\text{O}$ ,  $M_r = 18$ ]

Answer space continues on the next page.

5. (c) continued.

[for  $\text{Na}_2\text{SO}_4$ ,  $M_r = 142$

for  $\text{H}_2\text{O}$ ,  $M_r = 18$ ]

value of **X** = \_\_\_\_\_

(Total for Question 5 = 14 marks)

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**6. This question is about electrolysis.**

**(a) Look at the diagram for Question 6 (a) in the Diagram Booklet. The apparatus is used to collect the gases produced when an electric current passes through dilute sulfuric acid.**

**(i) Name a suitable material for the inert electrodes.  
(1 mark)**

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**(ii) Give a test to show that the gas produced at the positive electrode is oxygen.  
(1 mark)**

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**(continued on the next page)**

**6. (a) continued.**

- (iii) Give a test to show that the gas produced at the negative electrode is hydrogen.  
(1 mark)**

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- (iv) Give an ionic half-equation for the formation of hydrogen at the negative electrode.  
(1 mark)**

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6. (a) continued.

- (v) The oxygen gas in the tube has a volume of  $17.8 \text{ cm}^3$  at rtp.

Calculate the mass, in grams, of oxygen gas in the tube.

Give your answer to 3 significant figures.  
(3 marks)

[for  $\text{O}_2$  at rtp, molar volume =  $24\,000 \text{ cm}^3$ ]

mass of oxygen = \_\_\_\_\_ g

(continued on the next page)

**6. continued.**

**(b) Look at the diagram for Question 6 (b) in the Diagram Booklet. The apparatus is used to electrolyse a concentrated solution of sodium chloride.**

**(i) Give a reason why sodium chloride solution conducts electricity.  
(1 mark)**

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**(ii) The positive ions in the solution are  $\text{H}^+$  and  $\text{Na}^+$**

**Give a reason why hydrogen forms at the negative electrode rather than sodium.  
(1 mark)**

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**6. (b) continued.**

**(iii) A sample of the solution is taken from near the negative electrode.**

**The sample is tested with universal indicator.**

**Explain the final colour of the universal indicator.  
(2 marks)**

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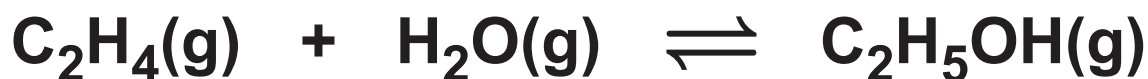
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**(Total for Question 6 = 11 marks)**

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7. Gaseous ethanol is manufactured by reacting ethene gas with steam.

This is the equation for the reaction.



- (a) (i) The pressure on an equilibrium mixture of the three gases is decreased. All the other conditions are kept the same.

Predict the effect of this change on the yield of ethanol at equilibrium, giving a reason for your answer.

(2 marks)

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**7. (a) continued.**

- (ii) Predict the effect of adding a catalyst on the yield of ethanol at equilibrium, giving a reason for your answer.  
(2 marks)**

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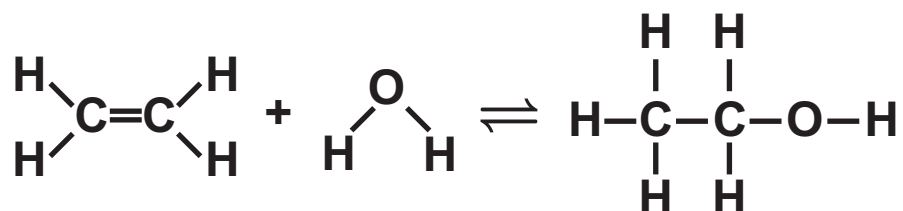
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7. continued.

- (b) The equation shows the displayed formulae for the reactants and products.



Look at the table for Question 7 (b) in the Diagram Booklet. It gives the bond energies.

- (i) Using the data in the table and the equation, show that the enthalpy change,  $\Delta H$ , is approximately  **$-40 \text{ kJ/mol}$**   
(3 marks)

Answer space continues on the next page.



**7. (b) (i) continued.**

**(continued on the next page)**

**7. (b) continued.**

- (ii) Explain, in terms of bonds broken and bonds made, why this reaction is exothermic.  
(2 marks)**

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7. (b) continued.

- (iii) Look at the diagram for Question 7 (b) (iii) in the Diagram Booklet. Draw an energy level diagram for the reaction between ethene and steam.

Label the enthalpy change,  $\Delta H$   
(3 marks)

- (c) Ethanoic acid reacts with ethanol to form ester **A** and water.

- (i) Look at the diagram for Question 7 (c) (i) in the Diagram Booklet. Complete the equation by adding the displayed formula of ester **A**.  
(2 marks)

- (ii) Give the name of ester **A**.  
(1 mark)

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(Total for Question 7 = 15 marks)

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**TOTAL FOR PAPER = 70 MARKS**  
**END OF PAPER**